

UNIT TEST GASES AROUND US YEAR 5

Study Guide

Unit Test Gases Around Us Year 5

Understanding the gases that surround us is an important part of learning about the environment, science, and our daily lives. In Year 5, students begin to explore the different gases present in the air and other places around us. This knowledge helps them appreciate the world better and understand how gases influence weather, health, and the planet's climate. This article provides a comprehensive overview of unit test gases around us for Year 5 students, explaining what gases are, their common types, where they are found, and why they matter.

What Are Gases?

Definition of Gases

Gases are one of the three main states of matter, alongside solids and liquids. Unlike solids and liquids, gases do not have a fixed shape or volume. Instead, they expand to fill any space available to them. Gases are made up of tiny particles called molecules that move freely and quickly.

Properties of Gases

Gases have unique features that distinguish them from solids and liquids:

- **Expandability:** Gases expand to fill their container.
- **Compressibility:** They can be compressed into a smaller space.
- **Low Density:** Gases are less dense than solids and liquids.
- **Mixability:** Different gases can mix easily without reacting.

Common Gases Around Us

Major Gases in the Air

The air we breathe is made up of several gases. Here are the most common:

- **Nitrogen (N₂):** Accounts for about 78% of the Earth's atmosphere.

- **Oxygen (O₂):** Makes up about 21%, essential for breathing and combustion.
- **Argon (Ar):** A noble gas, about 0.9%, used in lighting.
- **Carbon Dioxide (CO₂):** Around 0.04%, important for photosynthesis.
- **Trace Gases:** Other gases like neon, helium, methane, and hydrogen are present in tiny amounts.

Other Important Gases in the Environment

Apart from the air, gases are present in various other places:

- **Water Vapor:** Gaseous form of water, present in the air and responsible for humidity.
- **Natural Gases:** Such as methane, found in swamps, wetlands, and natural gas fields.
- **Pollutant Gases:** Such as sulfur dioxide (SO₂) and nitrogen oxides (NO_x), produced by burning fossil fuels.

Where Are Gases Found Around Us?

In the Atmosphere

The atmosphere is the layer of gases surrounding Earth. It plays a crucial role in protecting living organisms and regulating the climate.

In Our Homes and Schools

Gases are present in everyday life:

- **Cooking Gas:** Propane or butane used in stoves.
- **Air Conditioning:** Some air conditioners use gases like Freon.
- **Fire Safety:** Carbon monoxide, which is dangerous, can be produced by burning fuels without enough oxygen.

In Nature

Gases are found naturally in various environments:

- **Volcanoes:** Release gases like sulfur dioxide and carbon dioxide.
- **Swamps and Wetlands:** Emit methane gas.
- **Oceans:** Gases like oxygen dissolve into water, supporting marine life.

The Role of Gases in Our Environment

Supporting Life

Oxygen is essential for humans, animals, and plants. Without it, respiration would not be possible, and life as we know it would not exist.

Climate Regulation

Gases like carbon dioxide and methane help trap heat in the atmosphere, maintaining Earth's temperature. This natural process is called the greenhouse effect.

Pollution and Its Impact

Some gases released by human activities are harmful:

- **Air Pollution:** Gases like sulfur dioxide and nitrogen oxides cause smog and acid rain.
- **Global Warming:** Excess greenhouse gases increase Earth's temperature, leading to climate change.

How Do We Test for Gases?

Simple Tests at School

Students can perform basic experiments to identify gases:

- **Test for Oxygen:** Relight a glowing splint in a test tube; if it ignites brighter, oxygen is present.
- **Test for Carbon Dioxide:** Bubble gas through limewater; if it turns cloudy, CO₂ is present.
- **Test for Hydrogen:** Bring a lit splint near the gas; if it makes a popping sound, hydrogen is present.

Safety Precautions

Always remember to:

- Handle gases carefully, especially flammable or toxic ones.
- Perform experiments under supervision.
- Use proper protective equipment like goggles and gloves.

The Importance of Protecting Our Atmosphere

Why Should We Care?

Since gases are vital to life and the environment, it is important to keep our atmosphere clean and balanced. Pollution can harm living things and change climate patterns.

What Can We Do?

Students and everyone can contribute by:

- Reducing car and factory emissions.
- Planting trees to absorb carbon dioxide.
- Using energy wisely to decrease pollution.
- Recycling and disposing of waste properly.

Summary

Gases are all around us and play a vital role in supporting life, shaping our environment, and influencing weather and climate. The most common gases in the air include nitrogen, oxygen, argon, and carbon dioxide. Gases are found in many places, from the atmosphere to natural sources like volcanoes and wetlands. Understanding gases helps us appreciate the importance of protecting our environment and maintaining a healthy planet.

Remember, as Year 5 students, exploring the gases around us is the first step toward becoming environmentally conscious and scientifically curious! Keep asking questions, performing safe experiments, and learning more about the fascinating world of gases.

Unit Test Gases Around Us Year 5: An In-Depth Exploration

Understanding the gases present around us is fundamental to grasping the basics of chemistry, environmental science, and daily life. For students in Year 5, exploring the gases we encounter daily can be both fascinating and educational. This comprehensive guide aims to provide a detailed overview of the gases around us, focusing on their properties, sources, significance, and the importance of understanding them through the lens of scientific testing and observation.

Introduction to Gases in Our Environment

Gases are one of the three primary states of matter, alongside solids and liquids. They are characterized by their ability to expand to fill their container, their low density compared to solids and

liquids, and their invisibility in most cases. The gases surrounding us come from natural sources, human activities, and biological processes. Recognizing and understanding these gases is crucial for health, safety, and environmental conservation.

Common Gases Present Around Us

1. Nitrogen (N₂)

- Abundance: Constitutes about 78% of Earth's atmosphere.
- Properties: Colorless, odorless, inert under most conditions.
- Sources & Uses:
 - Makes up most of the air we breathe.
 - Used in food preservation, as a blanketing gas in packaging, and in the production of ammonia for fertilizers.
- Role in Life:
 - Essential for forming amino acids and proteins.
 - Humans cannot directly use atmospheric nitrogen; it must be fixed into usable forms.

2. Oxygen (O₂)

- Abundance: About 21% of the atmosphere.
- Properties: Colorless, odorless, vital for respiration.
- Sources & Uses:
 - Produced by plants during photosynthesis.
 - Used in hospitals, welding, and rocket propulsion.
- Importance:
 - Essential for breathing and cellular respiration in animals and humans.
 - Supports combustion.

3. Carbon Dioxide (CO₂)

- Abundance: Approximately 0.04% of the atmosphere.
- Properties: Colorless, odorless, slightly acidic.
- Sources:
 - Produced during respiration, combustion, and decay.
 - Emitted by volcanic activity.
- Uses & Significance:
 - Used in carbonated drinks.

- Plays a vital role in the Earth's carbon cycle and photosynthesis.
- Excessive CO₂ contributes to global warming.

4. Water Vapor (H₂O) in Gaseous Form

- Presence: Varies with humidity and temperature.
- Properties: Invisible in most cases; visible as clouds or fog.
- Sources:
 - Evaporation from water bodies.
 - Transpiration from plants.
- Importance:
 - Critical for weather patterns.
 - Facilitates the water cycle.

5. Other Trace Gases

- Argon, neon, helium, methane, and ozone are present in smaller quantities but are important in various contexts.
 - For example:
 - Helium in balloons.
 - Ozone in the upper atmosphere protecting us from UV rays.

Testing and Detecting Gases

Understanding gases often involves testing and detection, which helps in identifying their presence and properties. For Year 5 students, simple experiments and observations can introduce the concept of gas testing.

Common Gas Tests

- Oxygen Test:
 - Method: Insert a glowing splint into a test tube containing the gas.
 - Result: The splint reignites if oxygen is present.
- Carbon Dioxide Test:
 - Method: Bubble the gas through limewater (calcium hydroxide solution).
 - Result: Limewater turns cloudy or milky.
- Hydrogen Test:
 - Method: Place a lit splint at the mouth of a test tube containing the gas.

- Result: A 'pop' sound indicates hydrogen.
- Nitrogen Test:
 - More complex, but generally inert and does not support combustion or change limewater.

Importance of Gas Testing in Daily Life

- Detecting leaks in household gas supplies.
- Monitoring pollution levels.
- Ensuring safety in laboratories and industrial settings.

The Role of Gases in Daily Life and the Environment

Gases are woven into the fabric of our daily existence. Their presence influences health, safety, environment, and technology.

Health and Safety

- Air Quality:
 - Clean air with appropriate levels of oxygen.
 - Monitoring for harmful gases like carbon monoxide (CO), which is odorless and deadly.
- Household Gases:
 - Natural gas (methane) used for cooking and heating.
 - Importance of proper ventilation and gas detectors.

Environmental Impact

- Greenhouse Effect:
 - Gases like CO₂ and methane trap heat in the atmosphere, affecting global temperatures.
- Air Pollution:
 - Emissions from vehicles and factories release pollutants like nitrogen oxides and sulfur dioxide.
- Ozone Layer:
 - Ozone (O₃) in the upper atmosphere protects us from harmful UV radiation.
 - However, ground-level ozone causes smog and respiratory problems.

Technological Applications

- Welding:

- Using oxygen and acetylene gases.
- Medical Uses:
 - Oxygen therapy for patients.
- Recreation:
 - Helium in balloons.
 - Propane and butane in camping stoves.

Understanding the Importance of Gas Conservation and Environmental Awareness

As young learners, recognizing the importance of conserving gases and their impact on the environment is vital.

- Reduce, Reuse, Recycle:
 - Minimizing pollution reduces harmful gas emissions.
- Planting Trees:
 - Trees absorb CO₂, helping control greenhouse gases.
- Energy Conservation:
 - Using energy-efficient appliances reduces the burning of fossil fuels.
- Supporting Clean Energy:
 - Solar, wind, and hydroelectric power produce fewer harmful gases.

Conclusion: The Significance of Gases Around Us

Gases play an indispensable role in our world, from supporting life to influencing climate and enabling technology. Recognizing the properties, sources, and effects of various gases empowers us to appreciate the environment and take responsible actions. For Year 5 students, understanding gases through simple experiments, observations, and discussions builds a foundation for more advanced scientific learning and environmental consciousness.

By exploring the gases around us, we learn not only about the invisible substances that fill our atmosphere but also about our responsibility to preserve the delicate balance of our planet's environment. Whether it's the oxygen we breathe, the carbon dioxide plants need, or the pollutants we aim to reduce, gases are integral to life on Earth. Developing curiosity and knowledge about these gases encourages a lifelong interest in science, health, and environmental stewardship.

In summary, gases around us include vital components like nitrogen, oxygen, and carbon dioxide, each with unique properties and roles. Testing gases helps us understand their presence and effects, which is essential for health, safety, and environmental management. As young scientists, students should be encouraged to observe, experiment, and think critically about the gases in their

surroundings, fostering a responsible attitude towards science and the environment.

QuestionAnswer What are gases around us in our daily environment? Gases around us include air, which is made up of nitrogen, oxygen, carbon dioxide, and other gases that we breathe and that surround us all the time. Why is it important to understand gases for year 5 students? Understanding gases helps us learn about the air we breathe, how plants and animals depend on it, and how gases are used in everyday life, making us more aware of our environment. How can we observe gases around us without using special equipment? We can observe gases by watching their effects, like seeing steam from boiling water, feeling the wind, or noticing bubbles in fizzy drinks. What are some common gases used in everyday life? Common gases include oxygen (for breathing), carbon dioxide (used in fizzy drinks), and natural gas (used for cooking and heating). How do gases behave in the air around us? Gases spread out to fill any space evenly, are invisible, and can move quickly from one place to another, especially when there is wind or movement.

Related keywords: air quality, pollution, atmospheric gases, environmental science, chemistry, air testing, gas sensors, pollution monitoring, air composition, science experiments

Physical characteristics of gases answer key

Physical characteristics of gases answer key

Understanding the physical characteristics of gases is fundamental in the study of chemistry and physics. These properties help scientists and students alike to identify, compare, and predict the behavior of gases in different environments. An answer key to questions about these characteristics serves as an essential resource for students preparing for exams, teachers designing assessments, and professionals conducting experiments. This comprehensive guide explores the key physical properties of gases, their significance, and how they differ from liquids and solids.

Introduction to Physical Characteristics of Gases

Gases are one of the three primary states of matter, distinguished by their unique physical properties. Unlike solids and liquids, gases do not have a fixed shape or volume and are highly compressible. Their particles are widely spaced and in constant, random motion. The physical characteristics of gases include properties such as pressure, volume, temperature, density, and diffusion rates, which are described and explained through various gas laws.

Main Physical Characteristics of Gases

1. Shape and Volume

- Gases do not have a definite shape or volume. They tend to expand to fill any container they are placed in.
- The shape and volume of a gas are determined by the container's dimensions.
- This property makes gases highly adaptable and allows them to flow and diffuse easily.

2. Compressibility

- Gases are highly compressible due to the large spaces between particles.
- Applying pressure decreases the volume of a gas significantly.
- This property is exploited in various applications, such as in pneumatic systems and airbags.

3. Density

- Density of gases is defined as mass per unit volume (usually grams per liter, g/L).
- Gases generally have much lower densities compared to solids and liquids.
- Density can vary greatly with temperature and pressure.

4. Pressure

- Gas particles exert force upon the walls of their container, creating pressure.
- Pressure is defined as force per unit area (Pascals, Pa).
- It results from collisions of gas particles with container walls.

5. Temperature

- Temperature measures the average kinetic energy of gas particles.
- Increasing temperature increases particle velocity and kinetic energy.
- It influences other properties like pressure and volume, as described by gas laws.

6. Diffusion and Effusion

- Diffusion is the process by which gas particles spread from high to low concentration.
- Effusion is the passage of gas particles through a tiny hole without collisions.

- Both phenomena are faster in gases than in liquids or solids, due to particle spacing and movement.

Key Gas Laws Related to Physical Characteristics

1. Boyle's Law

- States that at constant temperature, the volume of a gas is inversely proportional to its pressure.
- Mathematically: $P_1V_1 = P_2V_2$
- Demonstrates the compressibility of gases.

2. Charles's Law

- States that at constant pressure, the volume of a gas is directly proportional to its temperature.
- Mathematically: $V_1/T_1 = V_2/T_2$
- Explains how gases expand when heated.

3. Gay-Lussac's Law

- States that at constant volume, the pressure of a gas is directly proportional to its temperature.
- Mathematically: $P_1/T_1 = P_2/T_2$
- Highlights the relationship between temperature and pressure.

4. Avogadro's Law

- States that equal volumes of gases at the same temperature and pressure contain equal numbers of molecules.
- Mathematically: $V/n = \text{constant}$
- Emphasizes the role of particle count in gas volume.

Additional Physical Properties of Gases

1. Molar Mass

- The mass of one mole of a gas, influencing its density and behavior.
- Gases with lower molar masses diffuse faster.

2. Speed of Gas Particles

- Particles move at different speeds depending on temperature and mass.
- Lighter gases tend to have higher average particle velocities.

3. Mean Free Path

- The average distance traveled by a particle before collision.
- Influences diffusion and effusion rates.

Answer Key to Common Questions on Physical Characteristics of Gases

- Q1: Why do gases have lower densities compared to solids and liquids?

A: Because gas particles are spread far apart, resulting in less mass per unit volume.

- Q2: How does increasing temperature affect gas pressure?

A: It increases the kinetic energy of particles, leading to more frequent and forceful collisions, thereby increasing pressure.

- Q3: Why are gases easily compressed?

A: Due to the large spaces between particles, applying pressure reduces the volume significantly.

- Q4: What is diffusion, and why does it happen faster in gases?

A: Diffusion is the spreading of particles from high to low concentration; it occurs faster in gases because of their high particle velocities and large spacing.

- Q5: How does molar mass affect the rate of diffusion?

A: Gases with lower molar mass diffuse faster because their particles move at higher average

velocities.

Practical Applications of the Physical Characteristics of Gases

Understanding these properties is crucial for numerous practical applications:

- Designing breathing apparatus and anesthesia delivery systems.
- Developing industrial processes involving gas compression and storage.
- Explaining atmospheric phenomena and weather patterns.
- Creating effusion and diffusion-based separation techniques, such as in gas chromatography.
- Engineering safety measures like pressure relief valves in pressurized containers.

Summary

The physical characteristics of gases are fundamental to understanding their behavior and interactions in various environments. Their lack of fixed shape and volume, high compressibility, low density, and ability to diffuse rapidly distinguish gases from solids and liquids. Gas laws like Boyle's, Charles's, Gay-Lussac's, and Avogadro's laws provide mathematical frameworks to describe these properties and predict how gases respond to changes in pressure, volume, and temperature. Recognizing these characteristics and their interrelationships is essential for students, educators, and professionals working with gases across scientific and industrial fields.

Conclusion

Mastering the physical characteristics of gases and understanding their answer key is essential for a comprehensive grasp of gas behavior. These properties not only form the foundation of theoretical chemistry and physics but also have practical implications across various technological and environmental applications. Whether studying for exams or applying this knowledge in real-world scenarios, a thorough understanding of gas properties enables better analysis, innovation, and problem-solving in science and engineering.

Note: For further study, consult textbooks or resources that provide detailed explanations, diagrams, and practice questions related to the physical characteristics of gases.

Physical Characteristics of Gases Answer Key

Understanding the physical characteristics of gases is fundamental to grasping their behavior, applications, and importance in various scientific and industrial processes. The properties of gases differ significantly from solids and liquids, owing to their unique molecular arrangements and interactions. The "answer key" to understanding these characteristics provides clarity and a

foundation for solving related problems in chemistry and physics. This article offers an in-depth review of the physical properties of gases, emphasizing key features, principles, and their practical implications.

Introduction to Gases and Their Physical Properties

Gases are one of the three primary states of matter, characterized by their ability to expand indefinitely, fill their containers completely, and exhibit highly compressible behavior. Unlike solids and liquids, gases have molecules that are widely spaced and move freely, leading to distinctive physical properties. These properties include shape, volume, density, pressure, temperature, and behavior under different conditions. Understanding these attributes is essential for applications ranging from meteorology to chemical engineering.

Key Physical Characteristics of Gases

The physical properties of gases can be summarized as follows:

- Shape and Volume
- Compressibility
- Density
- Pressure
- Temperature
- Viscosity
- Diffusion and Effusion

Let's explore each of these in detail.

Shape and Volume

Shape:

Gases do not have a fixed shape; instead, they assume the shape of their container. The molecules move randomly and rapidly, filling the entire space available.

Volume:

Gases have variable volume, which can be expanded or compressed easily. They do not have a definite volume unless confined.

Features:

- Gases are highly adaptable in shape and volume.
- Their particles are far apart, leading to negligible intermolecular forces that influence shape and volume.

Pros:

- Flexibility in filling containers of any shape or size.
- Useful in applications requiring uniform distribution, such as in pneumatic systems.

Cons:

- Difficult to contain or store without specialized equipment.
- Their behavior is sensitive to external conditions like temperature and pressure.

Compressibility

Gases are highly compressible due to the large distances between their molecules. When pressure is applied, molecules are forced closer together, reducing volume significantly.

Features:

- Compressibility is a defining characteristic, allowing gases to be stored under high pressure.
- This property is exploited in technologies such as compressed air tanks, scuba diving tanks, and refrigeration.

Pros:

- Enables efficient storage and transportation.
- Facilitates the use of gases in various mechanical and industrial processes.

Cons:

- High compressibility requires careful handling to prevent accidents.
- Changes in pressure can lead to significant volume changes, affecting system stability.

Density of Gases

Density is defined as mass per unit volume (g/L or kg/m³). Gases are generally less dense than solids and liquids because of their large intermolecular distances.

Features:

- Density varies with temperature and pressure.
- Gases become less dense when heated and more dense when compressed.

Factors Influencing Density:

- Increasing pressure increases density.
- Increasing temperature decreases density.

Pros:

- Low density makes gases suitable for lighter-than-air applications like balloons and airships.
- Easy to manipulate density through pressure and temperature adjustments.

Cons:

- Low density can lead to rapid dispersal, making containment challenging.
- Variability with environmental conditions can complicate precise measurements.

Pressure

Pressure is the force exerted by gas molecules per unit area on the walls of its container. It results from molecules colliding with container walls and other molecules.

Features:

- Measured in units such as atmospheres (atm), pascals (Pa), or torr.
- Directly proportional to temperature and number of molecules (according to the ideal gas law).

Practical Applications:

- Used in designing pressure vessels and engines.
- Critical in weather systems and atmospheric studies.

Pros:

- Adjustable through changes in volume and temperature.
- Fundamental in understanding gas behavior and laws.

Cons:

- Excessive pressure can be dangerous.
- Variations can lead to system failures if not properly managed.

Temperature and Its Effect on Gases

Temperature measures the average kinetic energy of molecules. It influences almost all other physical properties of gases.

Features:

- Increasing temperature increases molecular motion, leading to expansion and increased pressure when volume is constant.
- Decreasing temperature causes contraction and reduces pressure.

Relationship with Other Properties:

- Governed by Gay-Lussac's law and the ideal gas law.
- Affects density, pressure, and diffusion rates.

Pros:

- Temperature control allows precise manipulation of gas properties.
- Essential for processes like gas liquefaction.

Cons:

- High temperatures can lead to instability or decomposition.
- Precise temperature regulation can be technologically demanding.

Viscosity of Gases

Viscosity refers to a fluid's resistance to flow. In gases, it is relatively low compared to liquids but still significant in many contexts.

Features:

- Increases with temperature (unusual compared to liquids).
- Influenced by molecular size and shape.

Applications:

- Affects the flow of gases in pipelines and engines.
- Important in aerodynamics and atmospheric sciences.

Pros:

- Low viscosity facilitates efficient flow and mixing.
- Allows for smooth operation in many industrial applications.

Cons:

- Variability with temperature can influence flow rates unpredictably.
- Higher viscosity at lower temperatures can cause flow resistance issues.

Diffusion and Effusion

Diffusion:

The spontaneous mixing of gases due to molecular motion, leading to uniform distribution.

Effusion:

The process by which gas particles pass through tiny holes without collisions between particles.

Features:

- Rate of diffusion is influenced by molecular mass; lighter gases diffuse faster.
- Graham's law describes the rate of effusion.

Applications:

- Gas purification.
- Leak detection.

Pros:

- Useful in separating gases.
- Enables rapid mixing in industrial processes.

Cons:

- Diffusion is slow for heavier gases.
- Effusion requires small openings and precise control.

Real-World Implications of Gas Properties

Understanding the physical characteristics of gases is critical for practical applications:

- Weather and Climate:

The behavior of atmospheric gases influences weather patterns, cloud formation, and climate change.

- Industrial Processes:

Gas laws guide the design of reactors, storage tanks, and pipelines.

- Medical Applications:

Oxygen therapy and anesthesia depend on gas behavior under different conditions.

- Aerospace Engineering:

Aerodynamics and propulsion systems are grounded in gas dynamics.

- Environmental Science:

Monitoring pollutants and greenhouse gases relies on understanding gas diffusion and dispersion.

Conclusion

The physical characteristics of gases are fundamental to numerous scientific and practical applications. Their ability to expand, compress, diffuse, and respond to temperature and pressure changes makes them unique among states of matter. Recognizing the pros and cons of these properties allows scientists and engineers to harness gases effectively while mitigating inherent challenges. The "answer key" to understanding these characteristics lies in mastering the principles of gas laws, molecular behavior, and their environmental interactions. As technology advances, our comprehension of these properties continues to deepen, leading to more innovative and efficient uses of gases across diverse fields.

Question What are the main physical characteristics of gases? Gases are compressible, have indefinite shape and volume, and can expand to fill their containers uniformly. How does the density of gases compare to solids and liquids? Gases have much lower densities than solids and liquids because their particles are spread far apart. Why do gases have indefinite shape and volume? Because gas particles are free to move and are not held together tightly, gases do not have a fixed shape or volume. What is the effect of temperature on the physical characteristics of gases? Increasing temperature causes gases to expand and become less dense, while decreasing temperature causes contraction and increased density. How does pressure influence the physical properties of gases? Higher pressure compresses gases into smaller volumes, increasing their density, while lower pressure allows gases to expand. What is meant by the compressibility of gases? Compressibility refers to the ability of gases to decrease in volume significantly under pressure, unlike solids and liquids. How does the diffusion of gases demonstrate their physical characteristics? Gases diffuse quickly and uniformly throughout a container, showing their ability to spread and fill space efficiently. What is the relationship between the physical characteristics of gases and the gas laws? The physical characteristics such as volume, temperature, and pressure are related through gas laws like Boyle's, Charles's, and Avogadro's law. Why are gases considered to have indefinite shape and volume from a molecular perspective? Because gas molecules are widely spaced and move freely, they do not have fixed shape or volume, filling their containers

completely.

Related keywords: gas properties, gas behavior, physical properties of gases, gas laws, ideal gases, real gases, gas density, gas volume, gas pressure, gas temperature

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